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Ch 7 Chemistry

Study Guide

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Properties of  
the Elements~~

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Problems Organic

**Chemistry,**

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**McMurry, Alkenes-**

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Gr.10 Chemistry

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1

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Periodic Trends:

Electronegativit

y, Ionization

Energy, Atomic

Radius - TUTOR

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Compounds | How  
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Chapter 7

(Atomic  
Structure and  
Periodicity) -

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7: Section 1:

Ions ~~Ch-7~~

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Guide:

Chemistry. (7.1)

Ions Key Term.

To find the  
number of

valence

electrons in an  
atom of a

representative  
element, simply  
look at its



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group number.

(7.1) Ions Key

Term cont. A

positively

charged ion, or

cation, is

produced when an

atom loses one

or more valence

electrons. (7.1)

Ions Key Term

cont ...

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*Page 17/49*

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Chemistry. To  
find the number  
of valence  
electrons in an  
atom of a  
representative  
element, simply  
look at its  
group number. A  
positively

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charged ion, or  
cation, is  
produced when an  
atom loses one  
or more valence  
electrons.

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Guide. STUDY.  
PLAY. valence

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electrons. electrons in the highest occupied energy level of an element's atoms. Group Number, 1A= 1, 2A=2, 3A=3, 4A=4, 5A=5, 6A=6, 7A=7, 8A=8. How can you determine the number of valence

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Answers  
electrons an  
atom has from  
its position on  
the periodic  
table? Name by  
that

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PLAY. Anion. An ion with a negative charge.

The charge for an Anion is written as a number followed by a minus sign.

Cation. An ion with a positive charge. The charge for a cation is written as a



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number followed  
by a plus sign.  
Chemical  
formula.

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guide chapter 7.  
what follows an  
element's symbol  
in a chemical

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formula to

indicate the  
number of atoms  
of that element  
in one molecule  
of the compound?  
what does the  
chemical formula  
for an ionic  
compound  
represent?

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PLAY. Explain  
why most carbon  
compounds are  
classified as  
organic  
compounds.

Carbon compounds  
controls what  
happens in the

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cell of living organism and play key roles in the chemistry of living organisms. Why is lead no longer used in paints for plumbing pipes.

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Guide. Define

chemical bond.

Explain

formation of

cations and

anions. Draw

electron dot

diagrams. Write

electron

configurations

for ions. The

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force that holds  
two atoms  
together.

Cations form  
when an atom  
loses electrons.  
An anion forms  
when... Ex.

$K \rightarrow 1s^2 2s^2$   
 $2p^6 3s^2 3p^6$   
 $4s^1.$

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Vocabulary.  
Valence  
electron.  
Electron dot  
structure. Octet  
rule. Halide  
ion. the  
electrons in the

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Answers  
highest occupied  
energy level of  
an eleme...

diagrams that  
show valence  
electrons in the  
atoms of an  
elemen... in  
forming  
compounds, atoms  
tend to achieve  
the electron  
confi...

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Chemistry in  
Context Study  
Guide CHEM 109  
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Energy Storage  
either  
decreasing the  
hydrogen  
pressure or

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increasing the temperature, the H<sub>2</sub> can be re-released on demand. Metal-hydride storage systems are ideally suited for PEM fuel cells that require high-purity hydrogen. Because metal hydrides are

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selective and  
absorb only  
hydrogen and not  
larger gas  
molecules, they  
act  
simultaneously  
...

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Chemistry in  
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### Chemistry

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Guide. the force  
that holds two  
atoms together;  
may form by the  
attraction of a  
positive ion for  
a negative ion  
or by sharing  
electrons. a thr  
ee-dimensional  
geometric  
arrangement of

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particles in  
which each  
positive ion is  
surrounded by  
negative ions  
and each  
negative ion is  
surrounded by  
positive ions;  
vary in shape  
due to sizes and  
relative numbers  
of the ions  
bonded.

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Certain topics  
like the  
structure,  
synthesis, and  
nomenclature of  
alkyl halides  
have already



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been covered previously. This part of the course is primarily concerned with  $S_N1$  and  $S_N2$  reactions,  $E1$  and  $E2$  reactions, and alkene synthesis, structure, and nomenclature.

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The following is the suggested order of readings and practice problems.

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The AP®  
Chemistry exam  
is a 3-hour

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15-minute, end-of-course test comprised of 60 multiple-choice questions, for which you will have 1 hour and 30 minutes (this counts for 50% of your score) and 7 free-response questions, for which you will

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have 1 hour and  
45 minutes (this  
counts

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Chemistry. From

aluminum to

xenon, we

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explain the  
properties and  
composition of  
the substances  
that make up all  
matter.

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